

Chadwick D. Sofield,
Marc D. Walter* and
Richard A. AndersenChemical Sciences Division, Lawrence Berkeley
Laboratory, and Department of Chemistry,
University of California, Berkeley, CA 94720,
USA

Correspondence e-mail: mwalter@rhrk.uni-kl.de

Key indicators

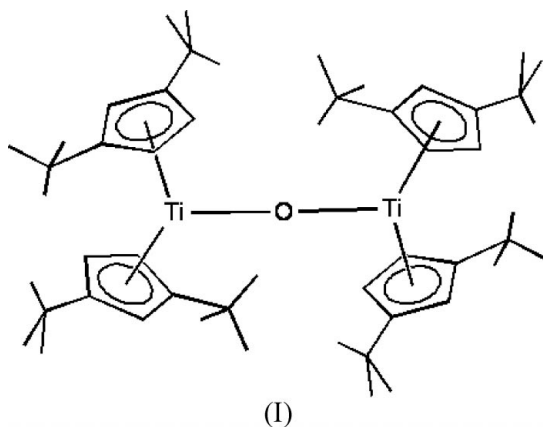
Single-crystal X-ray study
 $T = 164$ K
Mean $\sigma(\text{C}-\text{C}) = 0.006$ Å
 R factor = 0.065
 wR factor = 0.174
Data-to-parameter ratio = 14.6For details of how these key indicators were
automatically derived from the article, see
<http://journals.iucr.org/e>. μ_2 -Oxo-bis[bis(η^5 -1,3-di-*tert*-butylcyclopenta-
dienyl)titanium(III)]

The title compound, $[\text{Ti}_2\text{O}(\text{C}_{13}\text{H}_{21})_4]$ or $\{[(1,3\text{-}^t\text{Bu}_2\text{-C}_5\text{H}_3)_2\text{Ti}(\mu\text{-O})]\}$, crystallizes as a centrosymmetric molecule, in which the two metallocene fragments are bridged by an O atom. The trigonal planar coordination around each metal consists of two 1,3- $^t\text{Bu}_2\text{C}_5\text{H}_3$ ligands and the bridging O atom. The O atom is located on a crystallographic inversion center, resulting in a linear Ti—O—Ti linkage.

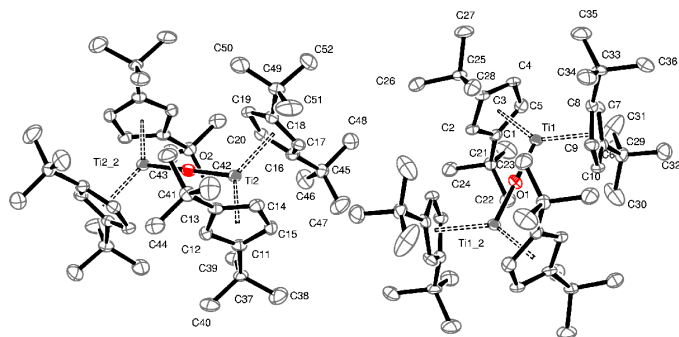
Received 27 August 2004
Accepted 6 September 2004
Online 11 September 2004

Comment

The title compound, (I), was prepared as part of an extensive study of titanium complexes containing the sterically demanding 1,3-di(*tert*-butyl)cyclopentadienyl (Cp^t) ligand. Although several μ -oxo complexes of Ti^{IV} are known, their Ti^{III} analogs, $\{[(\text{C}_5\text{H}_5\text{-}n\text{R}_n)_2\text{Ti}(\mu\text{-O})]\}$ (R is alkyl or aryl), are still rare, with the exceptions of the structurally characterized $\{[(\text{C}_5\text{H}_5)_2\text{Ti}(\mu\text{-O})]\}$ (Honold *et al.*, 1986), in which weak inter- and intramolecular ferromagnetic coupling has been observed (Lukens & Andersen, 1995), and the tetrahydroindenyl derivative, $\{[(\text{C}_9\text{H}_{11})_2\text{Ti}(\mu\text{-O})]\}$ (Kempe *et al.*, 1998).



Complex (I) crystallizes in the triclinic space group $P\bar{1}$ with two half-molecules in the asymmetric unit. The O atom of each molecule is located on an inversion center. The coordination around the Ti center is trigonal planar and the sum of the bond angles is 360° (the cyclopentadienyl ligands are represented by their centroids: $Cg1$, $Cg2$, $Cg3$ and $Cg4$ are the centroids of rings C1–C5, C6–C10, C11–C15 and C16–C20, respectively). A similar trigonal-planar geometry can be found in the dinuclear complexes $\{[(\text{C}_5\text{H}_5)_2\text{Ti}(\mu\text{-O})]\}$ (Honold *et al.*, 1986), $\{[(\text{C}_5\text{Me}_5)_2\text{Ti}(\mu\text{-N}_2)]\}$ (Sanner *et al.*, 1976), $\{[(\text{C}_5\text{HMe}_4)_2\text{Ti}(\mu\text{-N}_2)]\}$ (de Wolf *et al.*, 1996) and $\{[(1,3\text{-}(\text{SiMe}_3)_2\text{C}_5\text{H}_3)_2\text{Ti}(\mu\text{-N}_2)]\}$ (Hanna *et al.*, 2004). The average $Cg\text{-Ti-Cg}$ angle between the cyclopentadienyl ligands of $132.3(1)^\circ$ in (I) is slightly smaller than the value of 136° in the


Figure 1

A view of the two independent centrosymmetric molecules of (I), with displacement ellipsoids drawn at the 50% probability level. H atoms have been omitted for clarity.

C_5H_5 case (Honold *et al.*, 1986) and the value of 134.4° in the C_9H_{11} case (Kempe *et al.*, 1998).

The *tert*-butyl substituents on the cyclopentadienyl ligands adopt a common orientation in which the *tert*-butyl groups ‘gear’ in an alternating ‘backwards–forwards’ fashion with respect to the μ -oxo ligand. Similar conformations are observed in $\{[(1,3-(SiMe_3)_2C_5H_3)_2Ti]_2(\mu-N_2)\}$ (Hanna *et al.*, 2004) and $\{[(1,3-tBu_2C_5H_3)_2UO]_2\}$ (Lukens *et al.*, 1999).

The average Ti–O distance in (I) of 1.8769 \AA is on the upper limit ($1.83\text{--}1.86 \text{ \AA}$) expected for oxo-bridged dicyclopentadienyl–Ti complexes (Honold *et al.*, 1986, and references therein). A linear Ti–O–Ti linkage for this type of complex has not been observed to date. However, linearity has been observed in the Ti^{IV} complexes $\{[(C_5H_5)_2TiCl_2](\mu-O)\}$ (Thewalt & Schomburg, 1977) and $\{[benzyl_3Ti]_2(\mu-O)\}$ (Cetinkaya *et al.*, 1980). According to the model developed by Lauher & Hoffman (1976) for complexes of the type $(C_5H_5)_2ML_n$ ($n = 1, 2, 3$), the b_2 fragment orbital has suitable symmetry to interact with the p orbital of the oxygen bridge, forming a π bond. The orientation of the Cp' rings allows the empty b_2 symmetry orbital on each metallocene fragment to interact with a p orbital (either p_x or p_y) of the bridging O atom. In compound (I), the metallocene fragments are related by an inversion center, in contrast with $\{[(C_5H_5)_2Ti]_2(\mu-O)\}$ (Honold *et al.*, 1986) and $\{[(C_9H_{11})_2Ti]_2(\mu-O)\}$ (Kempe *et al.*, 1998), in which the metallocene fragments are rotated relative to each other by 90° and 63° , respectively, thereby maximizing the metallocene orbital overlap with that of O. This behavior is presumably due to the severe steric restrictions imposed on the molecule by the bulky *tert*-butyl groups.

Experimental

Compound (I) was synthesized by the reaction of base-free $[(1,3-tBu_2C_5H_3)_2Ti]$ (1.0 g, 2.5 mmol) in *n*-hexane with N_2O (1 atm) at room temperature. The color of the solution changed from blue to red–brown. After stirring the solution for 12 h, it was filtered, and crystallization from a saturated *n*-hexane solution at 253 K yielded red–brown crystals of (I) in 45% yield (m.p. 480–482 K). The crystal was handled under a nitrogen atmosphere, mounted on a glass fiber with Paratone-*N* oil and cooled to 164 (2) K for data collection.

Crystal data

$[Ti_2O(C_{13}H_{21})_4]$
 $M_r = 820.93$
 Triclinic, $P\bar{1}$
 $a = 10.6900$ (8) Å
 $b = 11.7225$ (8) Å
 $c = 20.2272$ (13) Å
 $\alpha = 90.196$ (2) $^\circ$
 $\beta = 101.945$ (2) $^\circ$
 $\gamma = 103.992$ (2) $^\circ$
 $V = 2402.5$ (3) Å^3

$Z = 2$
 $D_x = 1.135 \text{ Mg m}^{-3}$
 Mo $K\alpha$ radiation
 Cell parameters from 2932 reflections
 $\theta = 2\text{--}25^\circ$
 $\mu = 0.37 \text{ mm}^{-1}$
 $T = 164$ (2) K
 Block, red–brown
 $0.30 \times 0.25 \times 0.10 \text{ mm}$

Data collection

Bruker SMART CCD area-detector diffractometer
 ω scans
 Absorption correction: multi-scan (Blessing, 1995)
 $T_{\min} = 0.898$, $T_{\max} = 0.964$
 10 782 measured reflections
 7637 independent reflections
 4690 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.046$
 $\theta_{\max} = 25.6^\circ$
 $h = -12 \rightarrow 12$
 $k = -14 \rightarrow 12$
 $l = -20 \rightarrow 24$

Refinement

Refinement on F^2
 $R[F^2 > 2\sigma(F^2)] = 0.065$
 $wR(F^2) = 0.174$
 $S = 0.99$
 7637 reflections
 523 parameters
 H-atom parameters constrained
 $w = 1/[\sigma^2(F_o^2) + (0.0907P)^2]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} = 0.001$
 $\Delta\rho_{\max} = 1.24 \text{ e \AA}^{-3}$
 $\Delta\rho_{\min} = -0.40 \text{ e \AA}^{-3}$

Table 1

Selected geometric parameters (Å , $^\circ$).

Cg_1 , Cg_2 , Cg_3 and Cg_4 are the centroids of the rings C1–C5, C6–C10, C11–C15 and C16–C20, respectively.

Ti1–C1	2.470 (4)	Ti2–C11	2.485 (4)
Ti1–C2	2.456 (4)	Ti2–C12	2.455 (4)
Ti1–C3	2.479 (4)	Ti2–C13	2.450 (4)
Ti1–C4	2.418 (4)	Ti2–C14	2.407 (4)
Ti1–C5	2.416 (4)	Ti2–C15	2.421 (4)
Ti1–C6	2.499 (4)	Ti2–C16	2.514 (4)
Ti1–C7	2.516 (4)	Ti2–C17	2.523 (4)
Ti1–C8	2.495 (4)	Ti2–C18	2.516 (4)
Ti1–C9	2.396 (4)	Ti2–C19	2.398 (4)
Ti1–C10	2.393 (4)	Ti2–C20	2.391 (4)
Ti1–Cg1	2.134	Ti2–Cg3	2.159
Ti1–Cg2	2.146	Ti2–Cg4	2.128
Ti1–O1	1.8780 (8)	Ti2–O2	1.8760 (8)
<hr/>			
Ti1–O1–Ti1 ⁱ	180	Cg1–Ti1–Cg2	132.43
Ti2–O2–Ti2 ⁱⁱ	180	Cg3–Ti2–Cg4	132.23

Symmetry codes: (i) $-x, 1-y, 1-z$; (ii) $-x, 2-y, -z$.

Methyl H atoms were placed in calculated positions and refined using a riding model, with C–H distances of 0.98 \AA and $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{C})$. For each methyl group, the torsion angle which defines the orientation about the C–C bond was refined. The remaining methine H atoms were included in the model at calculated positions using a riding model, with C–H distances of 0.95 \AA and $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$. Two maxima of residual unassigned electron density greater than 1 e \AA^{-3} were observed at 1.83 \AA from atom C20 and at 1.76 \AA from atom C9. Due to poor crystal quality, diffraction was only observed up to $\theta_{\max} = 25.6^\circ$ with 0.85 data completeness, but the data were virtually complete up to $\theta = 23.0^\circ$.

Data collection: *SMART* (Bruker, 1999); cell refinement: *SAINT* (Bruker, 2002); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 1997); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997); molecular graphics: *ORTEP-3* (Farrugia, 1997); software used to prepare material for publication: *SHELXL97*.

This work was supported by the Director, Office of Energy Research, Office of Basic Energy Sciences, Chemical Sciences Division of the US Department of Energy under contract No. DE-AC03-76SF00098. The authors thank Dr Frederick J. Hollander (at CHEXRAY, the University of California at Berkeley X-ray diffraction facility) for assistance with the crystallography, and the state of Rhineland Palatinate (Germany) for a graduate grant (MDW).

References

- Blessing, R. H. (1995). *Acta Cryst.* **A51**, 33–38.
- Bruker (1999). *SMART*. Version 5.054d. Bruker AXS Inc., Madison, Wisconsin, USA.
- Bruker (2002). *SAINT*. Version 6.40. Bruker AXS Inc., Madison, Wisconsin, USA.
- Cetinkaya, B., Hitchcock, P. B., Lappert, M. F., Torroni, S., Atwood, W. E., Hunter, W. E. & Zoworotko, M. J. (1980). *J. Organomet. Chem.* **188**, C31–C35.
- Farrugia, L. J. (1997). *J. Appl. Cryst.* **30**, 565.
- Hanna, T. E., Keresztes, I., Lobkovsky, E., Bernskoetter, W. H. & Chirik, P. J. (2004). *Organometallics*, **23**, 3448–3458.
- Honold, B., Thewalt, U., Herberhold, M., Alt, H. G., Kool, L. B. & Rausch, M. D. (1986). *J. Organomet. Chem.* **314**, 105–111.
- Kempe, R., Spannenberg, A., Penlecke, N. & Rosenthal, U. (1998). *Z. Kristallogr. New Cryst. Struct.* **213**, 423–424.
- Lauher, J. W. & Hoffman, R. J. (1976). *J. Am. Chem. Soc.* **98**, 1729–1742.
- Lukens, W. W., Allen, P. G., Bucher, J. J., Edelstein, N. M., Hudson, E. A., Shuh, D. K., Reich, T. & Andersen, R. A. (1999). *Organometallics*, **18**, 1253–1258.
- Lukens, W. W. & Andersen, R. A. (1995). *Inorg. Chem.* **34**, 3440–3443.
- Sanner, R. D., Duggan, D. M., McKenzie, T. C., Marsh, R. E. & Bercaw, J. E. (1976). *J. Am. Chem. Soc.* **98**, 8358–8365.
- Sheldrick, G. M. (1997). *SHELXS97* and *SHELXL97*. University of Göttingen, Germany.
- Thewalt, U. & Schomburg, D. (1977). *J. Organomet. Chem.* **127**, 169–174.
- Wolf, J. J. de, Blaauw, R., Meetsma, A., Teuben, J. H., Gyepes, R., Varga, V., Mach, K., Veldman, N. & Spek, A. L. (1996). *Organometallics*, **15**, 4977–4983.